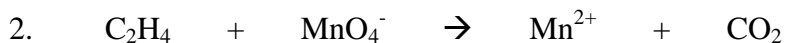


Balance the following redox reaction using the half-reaction method. Identify the oxidizing agent (OA), the reducing agent (RA), the element oxidized, and the element reduced.

**Acidic Solution Problems**

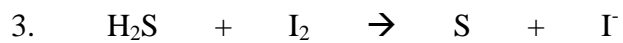
Element Oxidized \_\_\_\_\_ Element Reduced \_\_\_\_\_ RA \_\_\_\_\_ OA \_\_\_\_\_

Correct Equation:



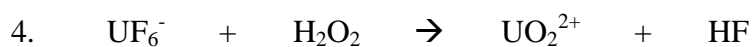
Element Oxidized \_\_\_\_\_ Element Reduced \_\_\_\_\_ RA \_\_\_\_\_ OA \_\_\_\_\_

Correct Equation:



Element Oxidized \_\_\_\_\_ Element Reduced \_\_\_\_\_ RA \_\_\_\_\_ OA \_\_\_\_\_

Correct Equation:



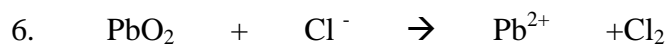
Element Oxidized \_\_\_\_\_ Element Reduced \_\_\_\_\_ RA \_\_\_\_\_ OA \_\_\_\_\_

Correct Equation:



Element Oxidized \_\_\_\_\_ Element Reduced \_\_\_\_\_ RA \_\_\_\_\_ OA \_\_\_\_\_

Correct Equation:



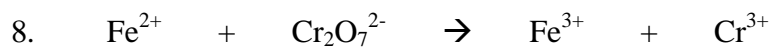
Element Oxidized \_\_\_\_\_ Element Reduced \_\_\_\_\_ RA \_\_\_\_\_ OA \_\_\_\_\_

Correct Equation:



Element Oxidized \_\_\_\_\_ Element Reduced \_\_\_\_\_ RA \_\_\_\_\_ OA \_\_\_\_\_

Correct Equation:



Element Oxidized \_\_\_\_\_ Element Reduced \_\_\_\_\_ RA \_\_\_\_\_ OA \_\_\_\_\_

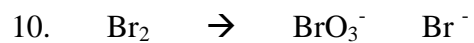
Correct Equation:



Element Oxidized \_\_\_\_\_ Element Reduced \_\_\_\_\_ RA \_\_\_\_\_ OA \_\_\_\_\_

Correct Equation:

### Basic Solution Problems



Element Oxidized \_\_\_\_\_ Element Reduced \_\_\_\_\_ RA \_\_\_\_\_ OA \_\_\_\_\_

Correct Equation:



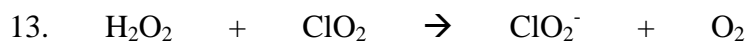
Element Oxidized \_\_\_\_\_ Element Reduced \_\_\_\_\_ RA \_\_\_\_\_ OA \_\_\_\_\_

Correct Equation:



Element Oxidized \_\_\_\_\_ Element Reduced \_\_\_\_\_ RA \_\_\_\_\_ OA \_\_\_\_\_

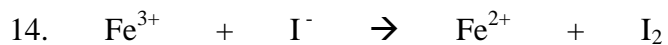
Correct Equation:



Element Oxidized \_\_\_\_\_ Element Reduced \_\_\_\_\_ RA \_\_\_\_\_ OA \_\_\_\_\_

Correct Equation:

### Neutral Solution Problems



Element Oxidized \_\_\_\_\_ Element Reduced \_\_\_\_\_ RA \_\_\_\_\_ OA \_\_\_\_\_

Correct Equation: