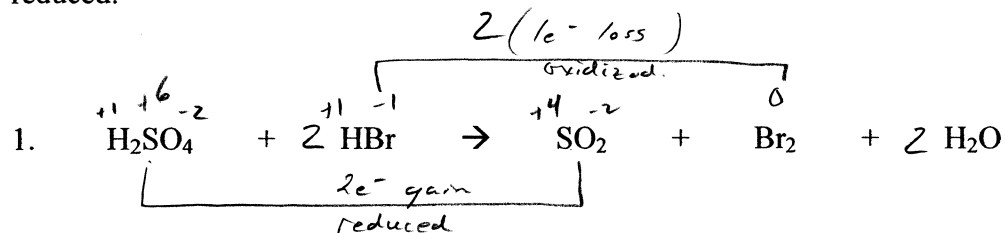
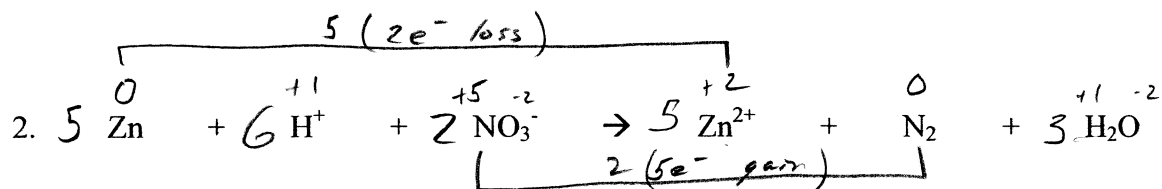


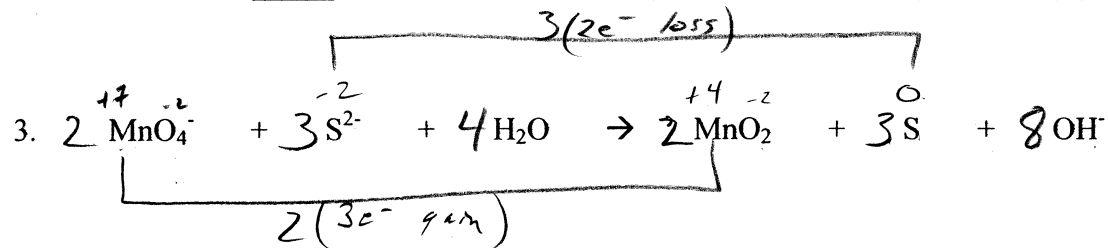
Balance the following redox reaction using the oxidation number method. Identify the oxidizing agent (OA), the reducing agent (RA), the element oxidized, and the element reduced.



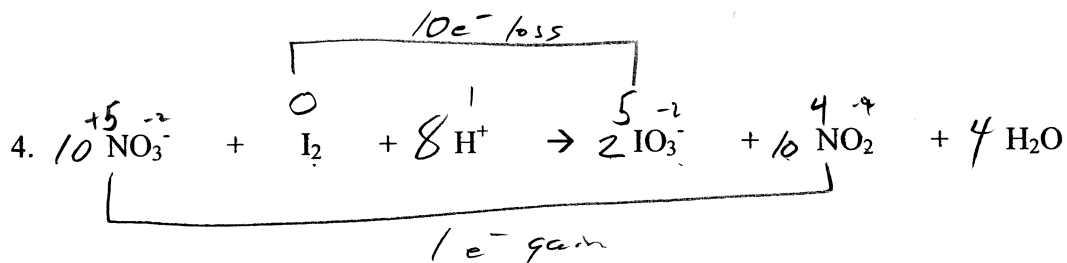
Element Oxidized Br Element Reduced S RA Br OA S



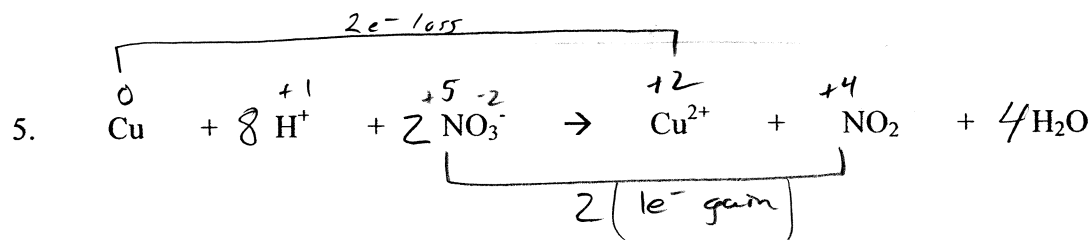
Element Oxidized Zn Element Reduced N RA Zn OA N



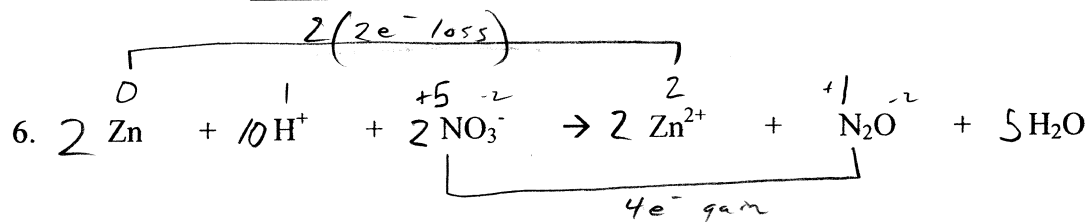
Element Oxidized S Element Reduced Mn RA S OA Mn



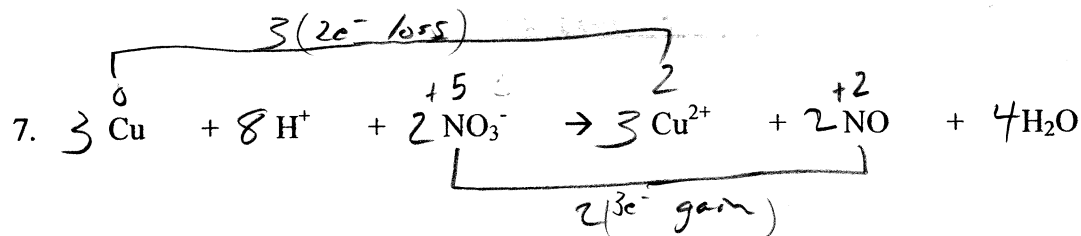
Element Oxidized I    Element Reduced N    RA I    OA N



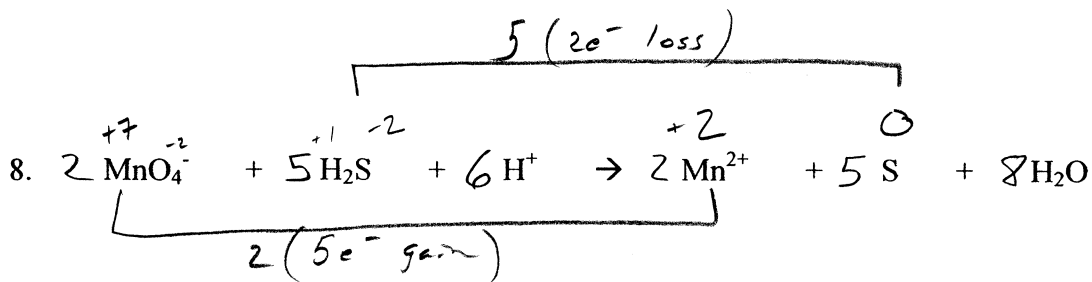
Element Oxidized Cu    Element Reduced N    RA Cu    OA N



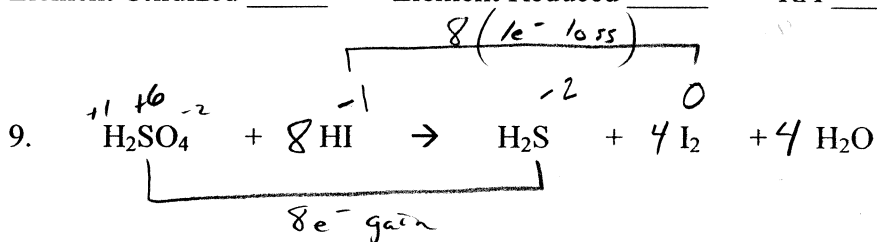
Element Oxidized Zn    Element Reduced N    RA Zn    OA N



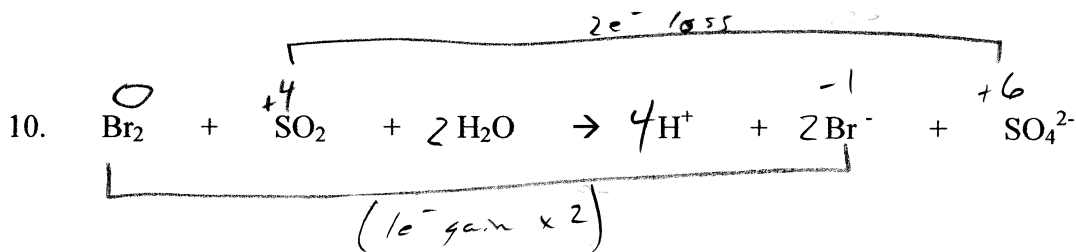
Element Oxidized Cu    Element Reduced N    RA Cu    OA N



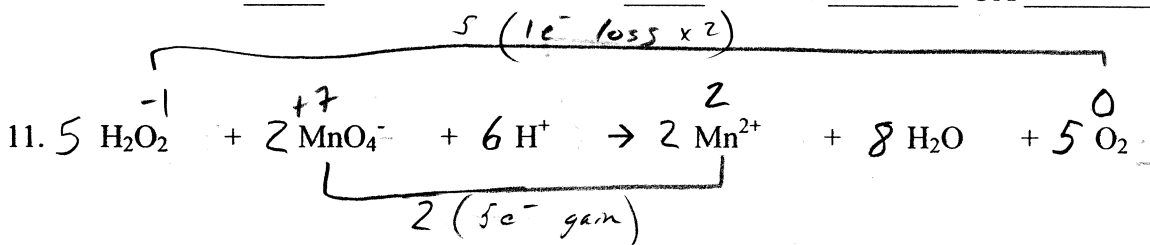
Element Oxidized S      Element Reduced Mn      RA S      OA Mn



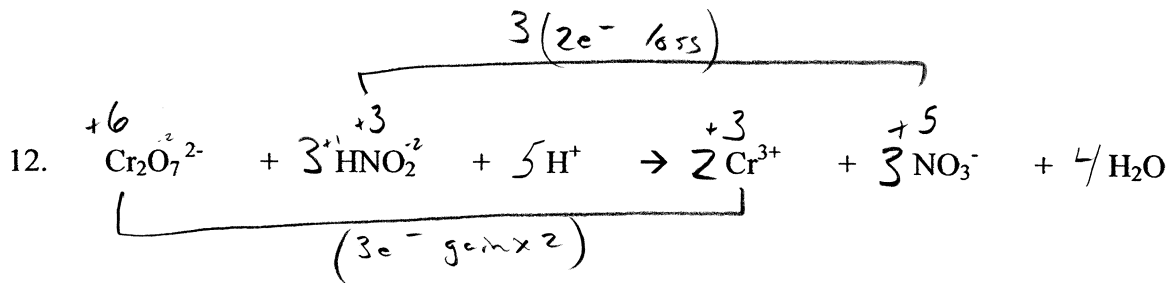
Element Oxidized I      Element Reduced S      RA I      OA S



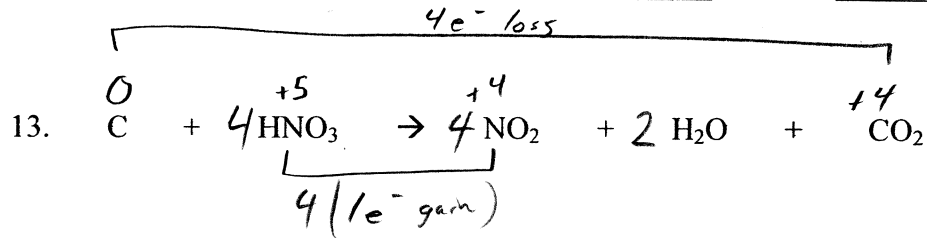
Element Oxidized S      Element Reduced Br      RA S      OA Br



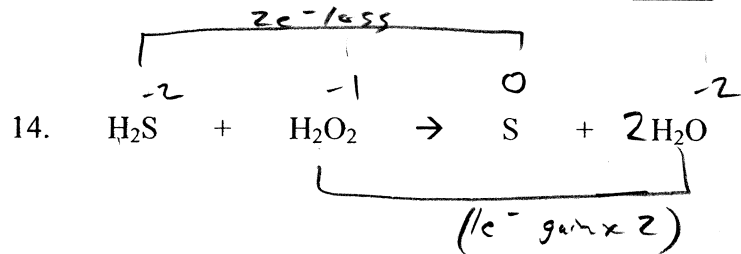
Element Oxidized O      Element Reduced Mn      RA O      OA Mn



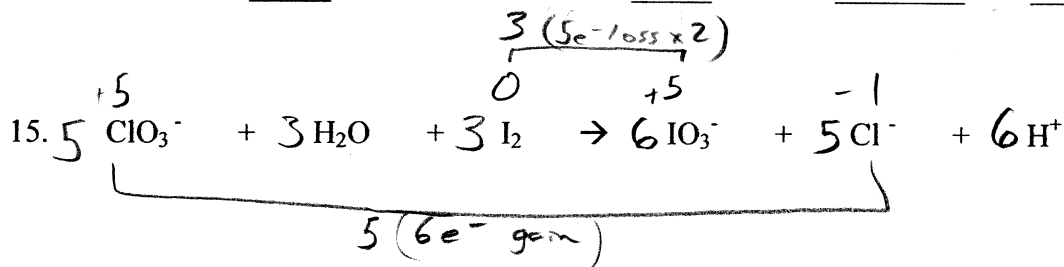
Element Oxidized N      Element Reduced Cr      RA N      OA Cr



Element Oxidized C      Element Reduced N      RA C      OA N



Element Oxidized S      Element Reduced O      RA S      OA O



Element Oxidized I      Element Reduced Cl      RA I      OA Cl