

## Electronegativities (BLM) KEY

Use the table of electronegativities to determine the bond type (ionic, polar covalent, nonpolar covalent) that would be formed between each of the following elements. Provide the electronegativity difference for each pair. Answers are based on Allred-Rochow table.

	<b>Predicted Bond Type</b>	<b>Electronegativity Difference</b>
1. Na, Cl	polar covalent	$2.83 - 1.01 = 1.72$
2. Al, Cl	polar covalent	$2.83 - 1.47 = 1.36$
3. H, S	ionic	$2.44 - 2.20 = 2.24$
4. K, F	ionic	$4.10 - 0.91 = 3.19$
5. O, O	non-polar covalent	$3.50 - 3.50 = 0$
6. Mg, S	polar covalent	$2.44 - 1.23 = 1.21$
7. Li, Br	polar covalent	$2.74 - 0.97 = 1.77$
8. F, F	non-polar covalent	$4.10 - 4.10 = 0$