

Practical Applications of Le Châtelier's Principle

There are many systems that rely on equilibrium conditions to work properly. When that system is stressed it returns to an equilibrium state according to Le Châtelier's principle.

Select two examples of such systems from the list given below and write a brief report describing:

- The equilibrium system (including chemical equations)
- Examples of possible stresses to the system
- Explain how the system returns to equilibrium with reference to Le Châtelier's principle

The report should be no more than one typed page using a 12-point font. **All sources must be documented. Reports will not be accepted unless sources are documented.**

The preferred documentation style is the ACS (American Chemical Society) style. Some documentation and examples are given here: <http://library.williams.edu/citing/styles/acs.php>

Examples:

Haber Process
haemoglobin production at high altitude
carbonated beverages
eyes adjusting to light
blood pH
recharging of batteries
turbocharged/supercharged engines
ester synthesis
weather indicators
carbonated beverages in a hen's diet