

p 861 36, 37, 38, 39, 40, 43, 45, 47, 50, 54, 56

$$\begin{aligned} \textcircled{36} \quad A_0 &= 1280 \text{ decays min}^{-1} \\ A &= 370 \text{ decays min}^{-1} \\ t &= 4.6 \text{ h} = 276 \text{ min} \end{aligned}$$

$$A = A_0 e^{-\lambda t}$$

$$\lambda = \frac{-\ln\left(\frac{A}{A_0}\right)}{t} = \frac{-\ln\left(\frac{370}{1280}\right)}{276 \text{ min}} = 0.005$$

$$T_{1/2} = \frac{\ln(2)}{\lambda} = \frac{\ln(2)}{0.005} = 138 \text{ min} = \underline{140 \text{ min}}$$

$$\textcircled{37} \quad (a) \quad \lambda = \frac{\ln(2)}{T_{1/2}} = \frac{\ln(2)}{4.5 \times 10^9 \text{ yr}} = \underline{1.5 \times 10^{-10} \text{ yr}^{-1}}$$

$$(b) \quad T_{1/2} = \frac{\ln(2)}{\lambda} = \frac{\ln(2)}{8.2 \times 10^{-5} \text{ s}^{-1}} = \underline{8500 \text{ s}}$$

$$\textcircled{38} \quad {}^{14}\text{C} \quad T_{1/2} = 5730 \text{ yr} \quad \lambda = \frac{\ln(2)}{T_{1/2}} = \frac{\ln(2)}{5730 \text{ yr}} = 1.2097 \times 10^{-4} \text{ yr}^{-1}$$

$$A = \lambda N = (1.2097 \times 10^{-4} \text{ yr}^{-1})(3.1 \times 10^{20}) = 3.75 \times 10^{16} \text{ decays yr}^{-1} \\ = \underline{1.19 \times 10^9 \text{ Bq}}$$

$$\textcircled{39} \quad \lambda = \frac{\ln(2)}{T_{1/2}} = \frac{\ln(2)}{(7/12) \text{ yr}} = 0.924 \text{ yr}^{-1}$$

$$\frac{A}{A_0} = e^{-\lambda t} = e^{-(0.924)(3)} = \underline{0.06}$$

$$(40) \quad \frac{A}{A_0} = e^{-\lambda t} = e^{-\frac{\ln 2}{T_{1/2}} (6 T_{1/2})} = \underline{0.02}$$

$$(43) \quad {}^{131}\text{I} \quad T_{1/2} = 8.0207 \text{ days} \quad \lambda = \frac{\ln 2}{T_{1/2}} = 1.0 \times 10^{-6} \text{ s}^{-1}$$

$$(a) \quad \frac{130.906124 \text{ g}}{682 \times 10^{-6} \text{ g}} = \frac{6.02 \times 10^{23}}{x}$$

$$N_0 = 3.136 \times 10^{18}$$

$$A_0 = \lambda N = (1.0 \times 10^{-6} \text{ s}^{-1}) (3.136 \times 10^{18}) = \underline{3.14 \times 10^{12} \text{ Bq}}$$

$$(b) \quad A = A_0 e^{-\lambda t} = (3.14 \times 10^{12} \text{ Bq}) e^{-(1.0 \times 10^{-6} \text{ s}^{-1})(3600 \text{ s})} = \underline{3.13 \times 10^{12} \text{ Bq}}$$

$$(c) \quad A = A_0 e^{-\lambda t} = (3.14 \times 10^{12} \text{ Bq}) e^{-(1.0 \times 10^{-6} \text{ s}^{-1})(1.578 \times 10^7 \text{ s})} = \underline{4.40 \times 10^5 \text{ Bq}}$$

$$(45) \quad \lambda = \frac{\ln 2}{T_{1/2}} = \frac{\ln 2}{(1.28 \times 10^9 \text{ yr})(3.156 \times 10^7 \text{ s yr}^{-1})} = 1.7153 \times 10^{-17} \text{ s}^{-1}$$

$$A = \lambda N$$

$$N = \frac{A}{\lambda} = \frac{2.0 \times 10^5 \text{ s}^{-1}}{1.7153 \times 10^{-17} \text{ s}^{-1}} = 1.166 \times 10^{22} \text{ nuclei}$$

$$\frac{1.166 \times 10^{22}}{6.02 \times 10^{23}} = 0.019 \text{ mol} (40 \text{ g mol}^{-1}) = \underline{0.77 \text{ g}}$$

$$(47) \quad \lambda = \frac{\ln 2}{T_{1/2}} = \frac{\ln 2}{7.55 \times 10^6 \text{ s}} = 9.181 \times 10^{-8} \text{ s}^{-1}$$

$$A = \lambda N$$

$$N = \frac{A}{\lambda} = \frac{2.65 \times 10^5 \text{ s}^{-1}}{9.181 \times 10^{-8} \text{ s}^{-1}} = 2.886 \times 10^{12} \text{ nuclei}$$

$$\frac{2.886 \times 10^{12}}{6.02 \times 10^{23}} = 4.795 \times 10^{-12} \text{ mol} (35 \text{ g mol}^{-1}) = \underline{1.68 \times 10^{-10} \text{ g}}$$

$$(50) \quad \begin{array}{l} 12.0107 \text{ g} = 6.02 \times 10^{23} \\ 285 \text{ g} \quad \quad \quad \times \\ 1.428 \times 10^{25} \text{ atoms} (1.3 \times 10^{-12}) = 1.857 \times 10^{13} \text{ } ^{14}\text{C atoms} \end{array}$$

$$T_{1/2} = 5730 \text{ yr} (3.156 \times 10^7 \text{ s yr}^{-1}) = 1.808 \times 10^{11} \text{ s}$$

$$\lambda = \frac{\ln(2)}{T_{1/2}} = \frac{\ln(2)}{1.808 \times 10^{11}} = 3.834 \times 10^{-12} \text{ s}^{-1}$$

$$A = \lambda N = (3.834 \times 10^{-12})(1.857 \times 10^{13}) = \underline{71.2 \text{ Bq}}$$

$$(54) \quad \begin{array}{l} T_{1/2} = 53 \text{ d} = 4.579 \times 10^6 \text{ s} \\ \lambda = \frac{\ln 2}{T_{1/2}} = 1.514 \times 10^{-7} \text{ s}^{-1} \end{array}$$

$$(a) \quad A = A_0 e^{-\lambda t}$$

$$t = \frac{-\ln\left(\frac{A}{A_0}\right)}{\lambda} = \frac{-\ln\left(\frac{15}{450}\right)}{1.514 \times 10^{-7}} = \underline{2.3 \times 10^7 \text{ s} = 4.3 \text{ days}}$$

54) (b) $A = \lambda N$

$$N = \frac{A}{\lambda} = \frac{450 \text{ s}^{-1}}{1.514 \times 10^{-7} \text{ s}^{-1}} = 2.97 \times 10^9 \text{ atoms}$$

$$\frac{2.97 \times 10^9}{6.02 \times 10^{23}} = 4.937 \times 10^{-15} \text{ mol} (7g) = \underline{3.5 \times 10^{-14} \text{ g}}$$

56) $\frac{12.0107 \text{ g}}{290 \text{ g}} = \frac{6.02 \times 10^{23}}{x}$
 $1.4535 \times 10^{25} \text{ atoms} (1.3 \times 10^{-12}) = 1.89 \times 10^{13} \text{ } ^{14}\text{C atoms}$

$$T_{1/2} = 5730 \text{ yr} = 1.808 \times 10^{11} \text{ s}$$

$$\lambda = \frac{\ln 2}{T_{1/2}} = 3.834 \times 10^{-12} \text{ s}^{-1}$$

$$A_0 = \lambda N_0 = (3.834 \times 10^{-12} \text{ s}^{-1})(1.89 \times 10^{13}) = 72.46 \text{ s}^{-1}$$

$$A = A_0 e^{-\lambda t}$$

$$t = \frac{-\ln\left(\frac{A}{A_0}\right)}{\lambda} = \frac{-\ln\left(\frac{8}{72.46}\right)}{3.834 \times 10^{-12}} = 5.748 \times 10^{11} \text{ s}$$

$$= \underline{18000 \text{ yrs}}$$